

Fill in the boxes on the following chart. You may use a Periodic chart to help you when necessary. Round Atomic weight to get the Mass Number and make sure you show all Neutron calculations.

$$\text{Atomic Number} = \text{Number of Protons} = \text{Number of Electrons}$$

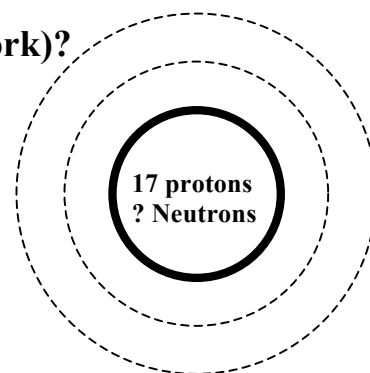
$$\text{Number of Protons} + \text{Number of Neutrons} = \text{Atomic Weight}$$

Name & Symbol	Atomic number	Atomic weight	# of Protons	# of Electrons	# of Neutrons	Mass Number
Strontium Sr	38	87.62		38		88
Hydrogen H	1	1.00				1
Vanadium V	23	50.94			$51-23 = 28$	51
Ruthenium Ru		101.07	44			101
Molybdenum Mo	42	95.94			$96-42 = 54$	96
Hafnium Hf		178.49	72			178
Europium Eu		151.96		63		152

Name & Symbol	Atomic number	Atomic weight	# of Protons	# of Electrons	# of Neutrons	Mass Number
Neodymium Nd		144.24	60			144
Krypton Kr	36	83.8		36		84
Indium In	49	114.82			$115-49 = 66$	115
Xenon Xe	54	131.3		54		131
Iridium Ir		192.22	77			192
Rubidium Rb		85.46		37		85
Terbium Tb	65	158.92	65		$159-65 = 94$	159

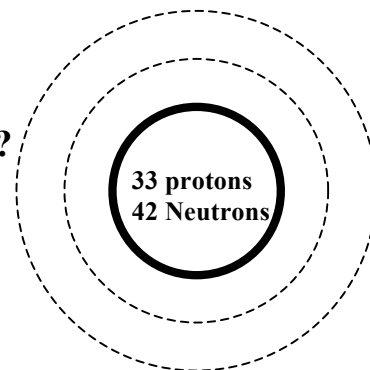
Name & Symbol	Atomic weight	# of Protons	# of Neutrons	# of Electrons	Mass Number	Atomic number
Bromine Br	79.90	35				35
Antimony Sb	121.75		$122-51 = 71$			
Helium He					4	2
Tungsten W	183.85	74				
Chromium Cr						
	196.96					
			$40-18 = 22$			
Tellurium Te					128	52
Cadmium Cd	112.41			48		
			$70-31 = 39$			31

1. How many neutrons does this atom have (show work)?
2. What is the atomic number for this atom?
3. How many electrons does this atom have?
4. What is the name of this atom?



Atomic Mass is 35

5. What is the atomic mass of this atom (show work)?
6. How many electrons does this atom have?
7. What is the name of this atom?



Atomic Mass is ?